Instructions: You have a total of 50 minutes to complete this test. Answer each of the following questions completely providing details and correct SI units.

Time Start_____ Time finish _____pledged _____

[1] An ideal 3 dimensional diatomic gas has the equation of state $PV=nRT$ where n is the number of moles and R=8.314 J/(K mol). You may assume n=1 mole here.

(a) Calculate c_p for this ideal gas when all 7 degrees of freedom are realized.

For the rest of this problem, assume $c_v=3/2\ R$.

(b) Suppose n=1 moles of this gas went through an isothermal process at T=300K which resulted in $V_f=3V_i$ with V_i being the initial volume of the gas. Calculate the following quantities:

$\Delta U=$_____

$Q=$_____

(c) Suppose n=1 moles of this gas went through an isobaric process at $P=1.0\text{x}10^2$ Pa and the volume changed from 1 m^3 to 3 m^3 . Calculate the following quantities:

$\Delta T=$_____

$\Delta U=$_____

$Q=$_____

[2] An unknown solid material is observed to have a length of 1 m at 0 ^{0}C and a length of 1.10 m at 100 ^{0}C.
(a) What is the coefficient of linear expansion for this material?

α=_____
(b) Another material (not the same material as in part a) has a coefficient of linear expansion given by $\alpha=2x10^{-3}$ /C. Calculate the coefficient of volume expansion for this material assuming the material is isotropic.

γ=_____
(c) If a cube of the material in (b) above with a volume $V_0=1$ m^3 is heated from 0 ^{0}C to 100 ^{0}C, calculate the work done when it expands against a constant pressure of $P=1x10^5$ Pa.

W=_____
(d) Suppose the material in (b) above has a mass of 800 kg and a specific heat of $c_p=2$ J/(kg ^{0}C). Calculate the heat supplied to the system.

Q=_____
(e) Calculate ΔU when the material in (b) above is expanded (as in c) and simultaneously heated (as in d).

ΔU=_____

[3] A heat engine is reported to operate with 25 % efficiency when the cold reservoir is at 0^0 C.

(a) Assuming this engine follows the Carnot cycle, what is the temperature of the hot reservoir?

T_H=_____

(b) Suppose the heat input to this engine was 10 J. Calculate the work done by this engine.

W=_____

(c) Suppose the heat input to this engine was 5J. Calculate the heat rejected by this engine.

Q_C=_____

[4] For water, $L_f=3.33x10^5$ J/kg and c=4186 J/Kg ^{0}C.
(a) If 4 kg of water at 0^0C is mixed with 1 kg of water at 80^0C, calculate the final equilibrium temperature of the mixture.

T_f=_____
(b) Calculate the change in entropy of the water as a result of the mixing described in part a. Hint: make sure your result is positive.

ΔS=_____
(c) If a 5 kg mass of ice at 0^0C melts to become water at 0^0C, calculate the change in entropy of the system.

ΔS=_____